

**Paper Reference(s) 1SC0/1CF**  
**Pearson Edexcel Level 1/Level 2**  
**GCSE (9–1)**

**Combined Science**  
**PAPER 2**  
**Foundation Tier**

**Diagram Booklet**

**In the boxes below, write your name, centre number and candidate number.**

<b>Surname</b>					
<b>Other names</b>					
<b>Centre Number</b>					
<b>Candidate Number</b>					

## **INSTRUCTIONS**

**There may be spare copies of some diagrams in case you need them.**

**THIS DIAGRAM BOOKLET MUST BE  
RETURNED WITH THE QUESTION PAPER  
AT THE END OF THE EXAMINATION.**

# **Contents**

## **Page**

<b>4</b>	<b>Question 1</b>
<b>5</b>	<b>Question 2(a)(i)</b>
<b>6</b>	<b>Question 2(a)(iv)</b>
<b>7</b>	<b>Question 2(b)</b>
<b>8</b>	<b>Question 3(c)(ii)</b>
<b>9</b>	<b>Question 4(c)</b>
<b>10</b>	<b>Question 4(c)(i)</b>
<b>11</b>	<b>Question 4(c)(i) (Spare copy)</b>
<b>12</b>	<b>Question 5(a)</b>
<b>13</b>	<b>Question 5(b)(iv)</b>
<b>14</b>	<b>Question 5(b)(iv) (Spare copy)</b>
<b>15</b>	<b>Question 5(c)</b>

## Question 1

### FIGURE 2

<b>substance</b>	<b>before heating</b>	<b>when hot</b>	<b>after cooling</b>
<b>metal spoon</b>	<b>solid</b>	<b>solid</b>	<b>solid</b>
<b>chocolate</b>	<b>solid</b>	<b>liquid</b>	<b>solid</b>
<b>egg white</b>	<b>liquid</b>	<b>solid</b>	<b>solid</b>

Question 2(a)(i)

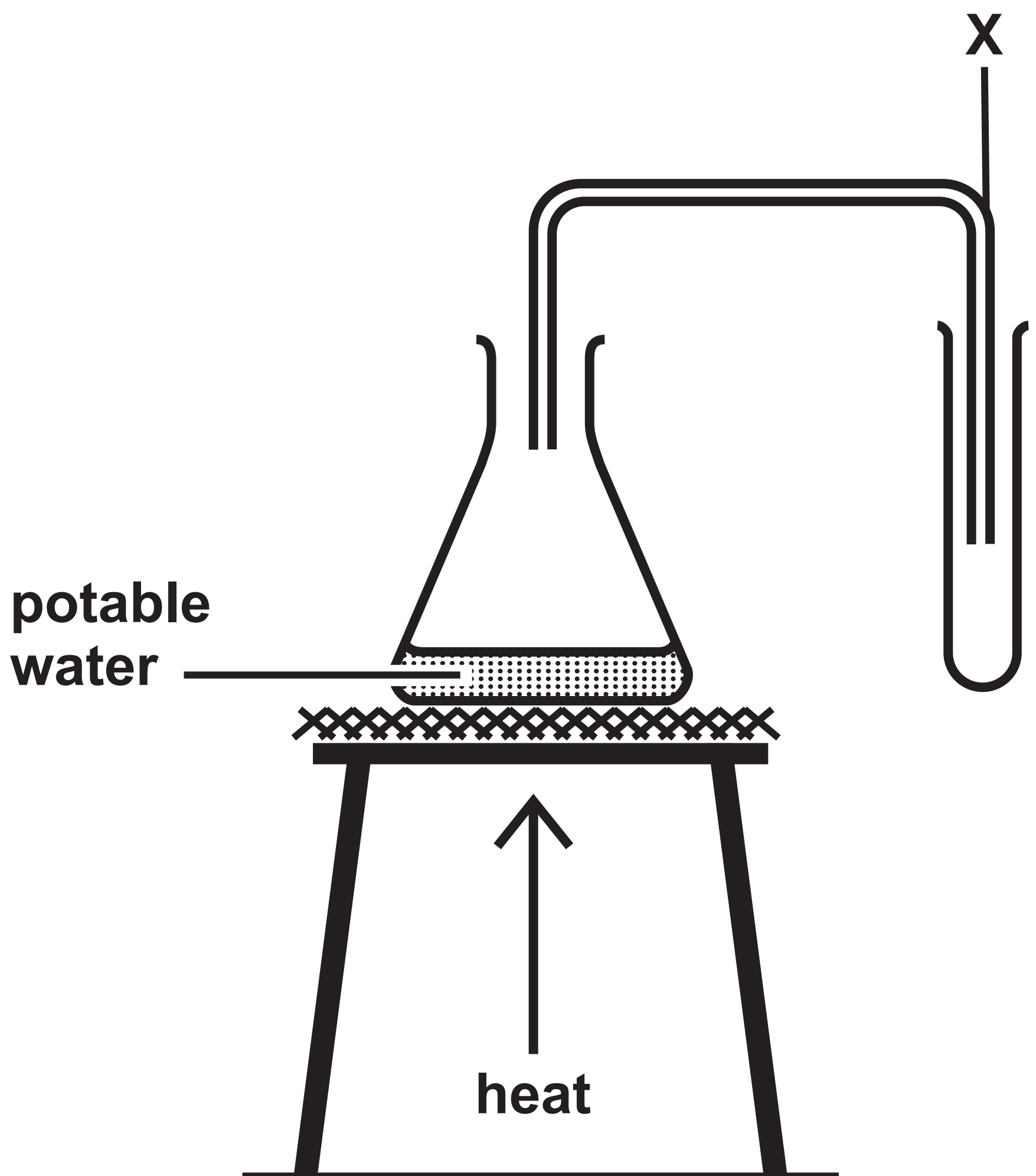
	first	second	third
<input type="checkbox"/> A	chlorination	sedimentation	filtration
<input type="checkbox"/> B	chlorination	filtration	sedimentation
<input type="checkbox"/> C	sedimentation	filtration	chlorination
<input type="checkbox"/> D	sedimentation	chlorination	filtration

**Question 2(a)(iv)****FIGURE 3**

<b>ion</b>	<b>concentration in mg dm<sup>-3</sup></b>
<b>chloride</b>	<b>60·70</b>
<b>fluoride</b>	<b>0·24</b>
<b>nitrate</b>	<b>24·90</b>
<b>sulfate</b>	<b>71·40</b>
<b>copper</b>	<b>0·05</b>
<b>magnesium</b>	<b>9·10</b>

## Question 2(b)

FIGURE 4



## Question 3(c)(ii)

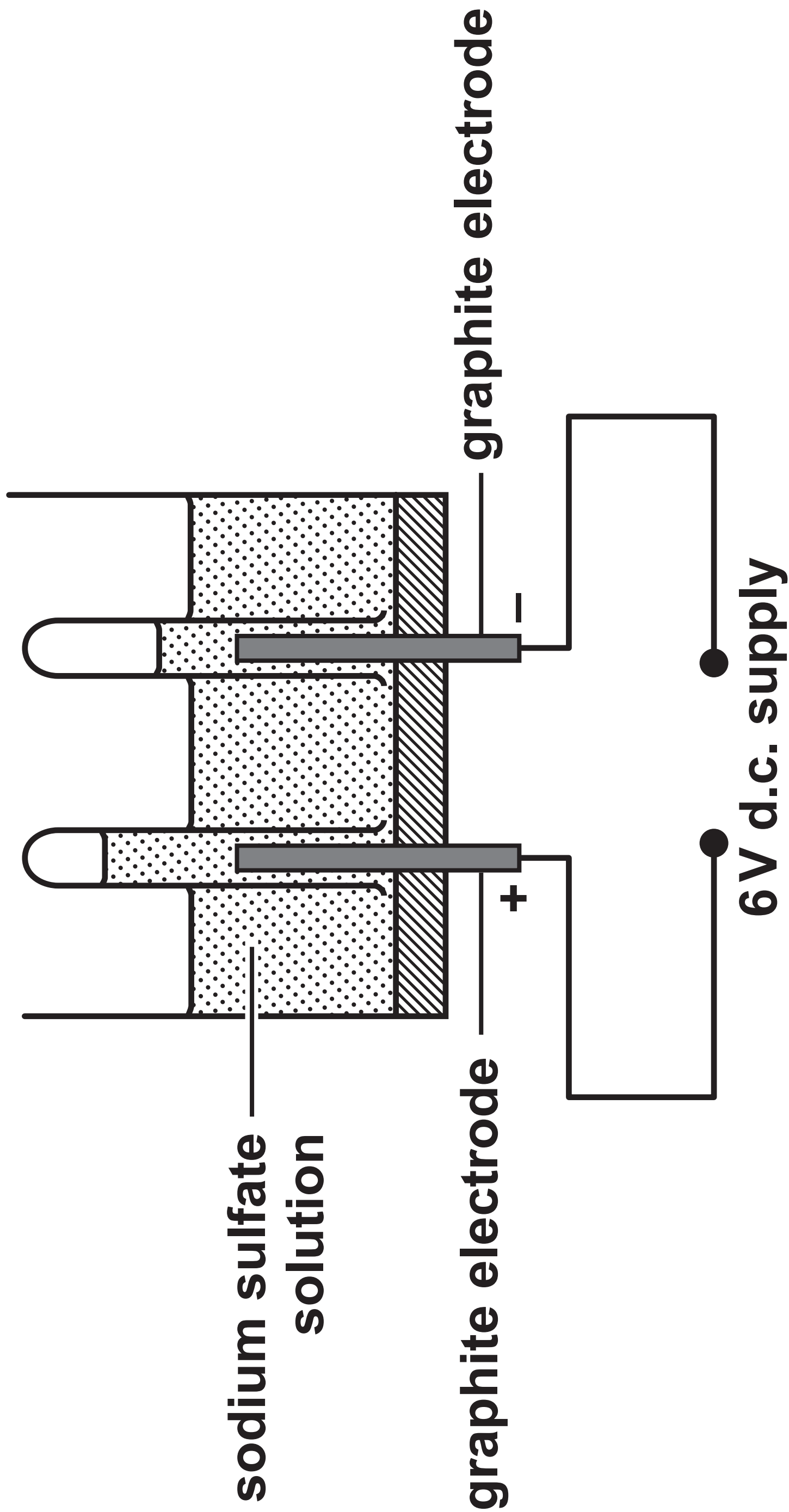
FIGURE 5

<b>property</b>	<b>predicted property</b>	<b>actual property</b>
<b>relative atomic mass</b>	<b>about 68</b>	<b>70</b>
<b>density in g/cm<sup>3</sup></b>	<b>about 6·0</b>	<b>5·9</b>
<b>melting point</b>	<b>lower than 40 °C</b>	<b>29·8 °C</b>
<b>density of oxide in g/cm<sup>3</sup></b>	<b>about 5·5</b>	<b>5·9</b>



Question 4(c)

FIGURE 6



## Question 4(c)(i)

## FIGURE 7



## Question 4(c)(i)

## FIGURE 7



Question 5(a)

barium hydroxide	hydrochloric acid	barium chloride	water
solid	aqueous	aqueous	liquid
solid	liquid	solid	aqueous
aqueous	aqueous	solid	liquid
aqueous	liquid	aqueous	aqueous

☐ A

☐ B

☐ C

☐ D

## Question 5(b)(iv)

pH of the  
mixture


mass of barium hydroxide in g

# Question 5(b)(iv)

pH of the  
mixture


mass of barium hydroxide in g

Question 5(c)

FIGURE 9

